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## **Chapter 11 Stoichiometry Study Guide**

CHAPTER 11: STOICHIOMETRY. UNIT 4: Chemical Reactions, The Mole, Stoichiometry and Thermodynamics. Part B:Stoichiometry. Big Picture Ideas: The identity of the reactants helps scientists to predict the products in a chemical reaction. Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum.

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TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 - Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles

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## **VIBRATIONS AND WAVES**

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368 Chapter 11 • Stoichiometry Section 11.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

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CHAPTER STUDY GUIDE Date Class Stoichiometry Section 11.1  
What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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CHAPTER Section 11.1 continued In your textbook, read about mole ratios. Answer the questions about the following chemical reaction. sodium + iron(III) oxide  $\rightarrow$  sodium oxide + iron  $6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$  15. What is a mole ratio? 16. How IS a mole ratio

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written? CA S Q C CYA 17. Predict the number of mole ratios for this reaction. Class 18.

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In this video, we will apply stoichiometric concepts to balanced equations in order to determine mass relationships within reactions.

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